



Suva®
refrigerants

**Thermodynamic
Properties
of
Suva® 407C
Refrigerant**

[R-407C (23/25/52)]

Thermodynamic Properties of Suva® 407C Refrigerant

SI Units

New tables of the thermodynamic properties of Suva® 407C refrigerant [ASHRAE designation: R-407C (23/25/52)], a near azeotropic blend of HFC-32/HFC-125/HFC-134a, have been developed and are presented here. These tables are based on extensive experimental measurements. Equations have been developed, based on the Peng-Robinson-Stryjek-Vera (PRSV) equation of state, which represent the data with accuracy and consistency throughout the entire range of temperature, pressure, and density presented in these tables.

Physical Properties

Chemical Formula	CH ₂ F ₂ /CHF ₂ CF ₃ /CH ₂ FCF ₃ (23/25/52% by weight)	
Molecular Weight	86.20	
Boiling Point at One Atmosphere	-43.56°C	(-46.40°F)
Critical Temperature, T _c	86.74°C 359.89 K	(188.13°F) (647.80°R)
Critical Pressure, P _c	4619.10 kPa (abs)	(669.95 psia)
Critical Density, D _c	527.30 kg/m ³	(32.92 lb/ft ³)
Critical Volume, V _c	0.00190 m ³ /kg	(0.0304 ft ³ /lb)

Units and Factors

t = temperature in °C	
T = temperature in K = °C + 273.15	
p _f = pressure of saturated liquid (bubble point) in kPa (abs)	
p _g = pressure of saturated vapor (dew point) in kPa (abs)	
v _f = volume of saturated liquid in m ³ /kg	
v _g = volume of saturated vapor in m ³ /kg	
V = volume of superheated vapor in m ³ /kg	
d _f = 1/v _f = density of saturated liquid in kg/m ³	
d _g = 1/v _g = density of saturated vapor in kg/m ³	
h _f = enthalpy of saturated liquid in kJ/kg	
h _{fg} = enthalpy of vaporization in kJ/kg	
h _g = enthalpy of saturated vapor in kJ/kg	
H = enthalpy of superheated vapor in kJ/kg	
s _f = entropy of saturated liquid in kJ/(kg) (K)	
s _g = entropy of saturated vapor in kJ/(kg) (K)	
S = entropy of superheated vapor in kJ/(kg) (K)	
C _p = heat capacity at constant pressure in kJ/(kg) (K)	
C _v = heat capacity at constant volume in kJ/(kg) (K)	

The gas constant, R = 8.314 J/(mole) (K)
for Suva® 407C, R = 0.09645 kJ/kg • K

One atmosphere = 101.325 kPa

Reference point for enthalpy and entropy:

h_f = 200 kJ/kg at 0°C

s_f = 1 kJ/kg • K at 0°C

Equations

The Peng-Robinson-Stryjek-Vera (PRSV) equation of state was used to calculate the tables of thermodynamic properties. It was chosen as the preferred equation of state because it provided an accurate fit of the thermodynamic data over the entire range of temperatures and pressures presented in these tables.

The constants for the PRSV equation of state were calculated in SI units. For conversion of thermodynamic properties to English (I/P) units, conversion factors are provided for each property derived from the PRSV equation of state.

1. Equation of State (PRSV)

$$P = RT/(V - b) - a/(V^2 + 2bV - b^2)$$

where P is in kPa, T is in K, V is in m³/mole, and R = 0.008314 kJ/(mole) (K). The constants a and b are calculated as follows:

$$a = \sum_{i=1}^3 \sum_{j=1}^3 x_i x_j a_{ij} \quad b = \sum_{i=1}^3 x_i b_i$$

where

$$a_{ij} = (a_i a_j)^{0.5} (1 - k_{ij}) \quad b_i = 0.077796 RT_{ci}/P_{ci}$$

x_i = mole fraction of component i

x_j = mole fraction of component j

$$a_i = (0.457235 R^2 T_{ci}^{-2}/P_{ci}) \alpha_i$$

$$a_j = (0.457235 R^2 T_{cj}^{-2}/P_{cj}) \alpha_j$$

k_{ij} = binary interaction parameter for components i and j

$$\alpha_i = [1 + \kappa_i (1 - T_{ri}^{0.5})]^2$$

$$\kappa_i = \kappa_{0i} + \kappa_{1i} [(1 + T_{ri}^{0.5}) (0.7 - T_{ri})]$$

(Note: $\kappa_i = \kappa_{0i}$ for $T_r > 0.7$)

$$\kappa_{0i} = 0.378893 + 1.4897153\omega_i - 0.17131848\omega_i^2 + 0.019655\omega_i^3$$

κ_{1i} = adjustable parameter for component i

$$T_{ri} = T_i/T_{ci}$$
 for component i

Values for R, T_{ci}, P_{ci}, ω_i, κ_{1i}, x_i, and k_{ij} are needed to calculate constants a and b. R = 0.008314 kJ/(mole) (K). The remaining constants for Suva® 407C are summarized below:

Component	T _{ci}	P _{ci}	ω _i	κ _{1i}	x _i
HFC-32 (i = 1)	351.60	5830.0	0.2763	-0.0250	0.38110
HFC-125 (i = 2)	339.19	3595.0	0.3023	0.0310	0.17956
HFC-134a (i = 3)	374.20	4056.0	0.3266	-0.0060	0.43934

The binary interaction parameters, k_{ij} , for Suva® 407C are:

$$\begin{array}{lll} k_{11} = 0.00000 & k_{12} = -0.00028 & k_{13} = -0.00815 \\ k_{21} = -0.00028 & k_{22} = 0.00000 & k_{23} = -0.00240 \\ k_{31} = -0.00815 & k_{32} = -0.00240 & k_{33} = 0.00000 \end{array}$$

Ideal Gas Heat Capacity Equation (at constant pressure):

$$C_p^o(\text{mixture}) = \sum_{i=1}^3 x_i C_{pi}^o$$

$$C_{pi}^o = 4.184 (A_i + B_i T + C_i T^2 + D_i T^3 + E_i T^4 + F_i T^5)$$

where C_p^o and C_{pi}^o are in J/(mole) (K) and T is in K.

x_i is the mole fraction of component i in the mixture (use same values listed in PRSV constants for Suva® 407C). A_i , B_i , C_i , D_i , E_i , and F_i are constants:

$$A_1 = 1.226880 \text{ E+01} \quad B_1 = -0.699113 \text{ E-01}$$

$$A_2 = 1.170140 \text{ E+01} \quad B_2 = 0.216411 \text{ E-01}$$

$$A_3 = 0.463685 \text{ E+01} \quad B_3 = 0.617904 \text{ E-01}$$

$$C_1 = 0.394642 \text{ E-03} \quad D_1 = -0.837462 \text{ E-06}$$

$$C_2 = 0.868526 \text{ E-04} \quad D_2 = -0.112776 \text{ E-06}$$

$$C_3 = -0.309907 \text{ E-04} \quad D_3 = 0.000000 \text{ E+00}$$

$$E_1 = 0.859548 \text{ E-09} \quad F_1 = 0.000000 \text{ E+00}$$

$$E_2 = 0.000000 \text{ E+00} \quad F_2 = 0.000000 \text{ E+00}$$

$$E_3 = 0.000000 \text{ E+00} \quad F_3 = 0.000000 \text{ E+00}$$

Properties calculated in SI units from the equations and constants listed above can be converted to I/P units using the conversion factors shown below. Please note that in converting enthalpy and entropy from SI to I/P units, a change in reference states must be included (from H = 200 and S = 1 at 0°C for SI units to H = 0 and S = 0 at -40°F for I/P units). In the conversion equations below, H (ref) and S (ref) are the saturated liquid enthalpy and entropy at -40°C. For Suva® 407C, H (ref) = 146.6 kJ/kg and S (ref) = 0.7903 kJ/kg • K.

Conversion Factors (SI units to I/P units):

$$P(\text{psia}) = P(\text{kPa}) \cdot 0.14504$$

$$T(\text{°F}) = (T(\text{°C}) \cdot 1.8) + 32$$

$$D(\text{lb/ft}^3) = D(\text{kg/m}^3) \cdot 0.062428$$

$$V(\text{ft}^3/\text{lb}) = V(\text{m}^3/\text{kg}) \cdot 16.018$$

$$H(\text{Btu/lb}) = [H(\text{kJ/kg}) - H(\text{ref})] \cdot 0.43021$$

$$S(\text{Btu/lb} \cdot \text{°R}) = [S(\text{kJ/kg} \cdot \text{K}) - S(\text{ref})] \cdot 0.23901$$

$$C_p(\text{Btu/lb} \cdot \text{°F}) = C_p(\text{kJ/kg} \cdot \text{K}) \cdot 0.23901$$

$$C_v(\text{Btu/lb} \cdot \text{°F}) = C_v(\text{kJ/kg} \cdot \text{K}) \cdot 0.23901$$

2. Vapor Pressure

$$\log_n P_{\text{sat}} = A + B/T + C \log_n T + D T^2$$

For SI units

T is in K and P is in kPa (abs)

A, B, C, and D are constants.

Constants for vapor pressure of saturated liquid (bubble point), p_f :

$$A = 4.27103 \text{ E+01} \quad C = -4.39387 \text{ E+00}$$

$$B = -3.34460 \text{ E+03} \quad D = 6.86997 \text{ E-06}$$

Constants for vapor pressure of saturated liquid (dew point), p_g :

$$A = 7.46912 \text{ E+01} \quad C = -9.51789 \text{ E+00}$$

$$B = -4.50059 \text{ E+03} \quad D = 1.73528 \text{ E-05}$$

For I/P units

T is in °R and P is in psia

A, B, C, and D are constants.

Constants for vapor pressure of saturated liquid (bubble point), p_f :

$$A = 4.33622 \text{ E+01} \quad C = -4.39387 \text{ E+00}$$

$$B = -6.02028 \text{ E+03} \quad D = 2.12036 \text{ E-06}$$

Constants for vapor pressure of saturated liquid (dew point), p_g :

$$A = 7.83549 \text{ E+01} \quad C = -9.51789 \text{ E+00}$$

$$B = -8.10106 \text{ E+03} \quad D = 0.53558 \text{ E-05}$$

3. Density of the Saturated Liquid

$$d_f/D_c = a_0 + a_1 z + a_2 z^2 + a_3 z^3 + a_4 z^4$$

$$\text{where } z = (1 - T/T_c)^{1/3} - t_0$$

Because both density and temperature appear in the reduced form in the equation, the same constants can be used for either SI or I/P units.

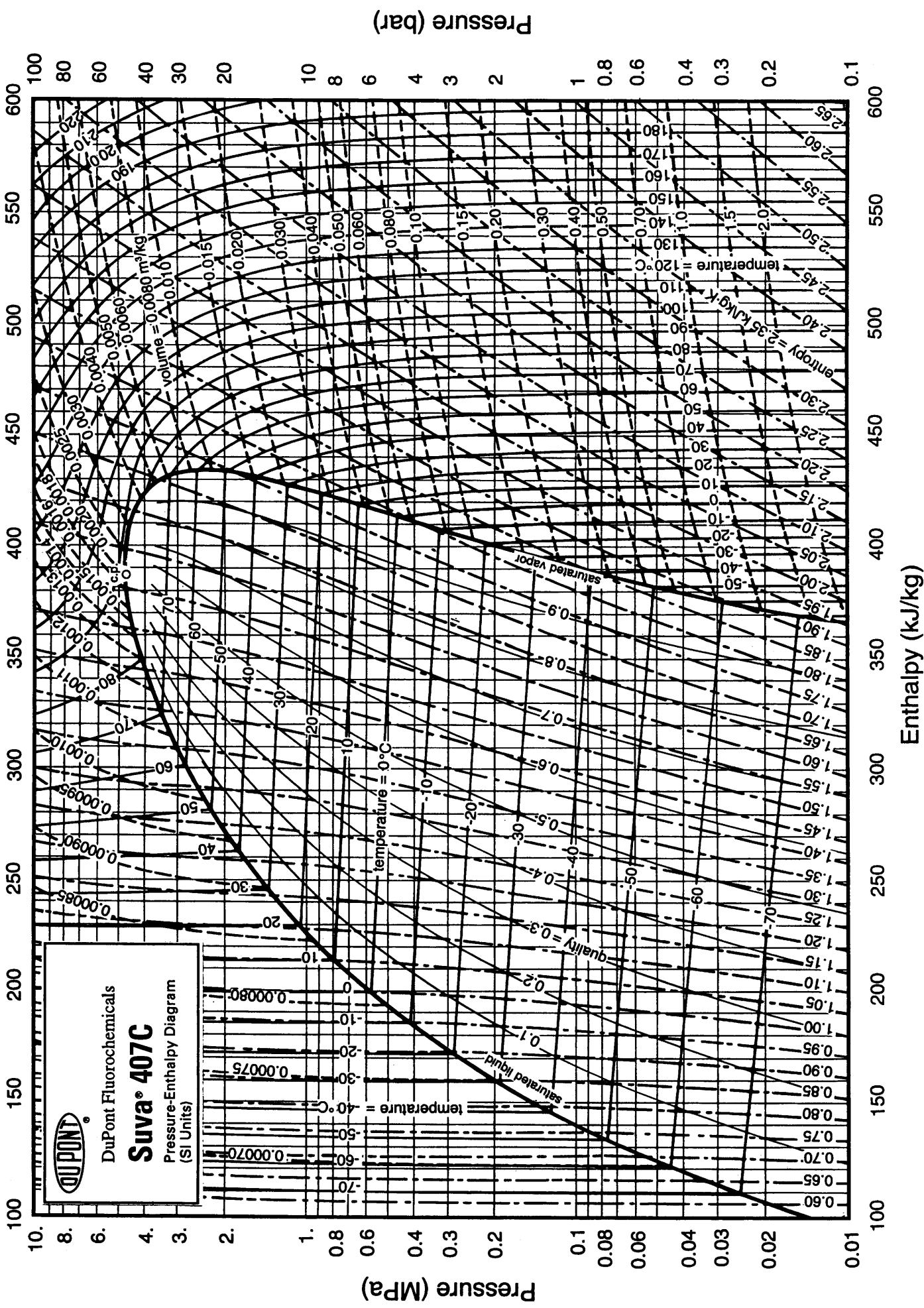
d_f and D_c are in kg/m³ in SI units and lb/ft³ in I/P units. T and T_c are in K in SI units and °R in I/P units.

a_0 , a_1 , a_2 , a_3 , a_4 , and t_0 are constants:

$$a_0 = 1.000000 \text{ E+00} \quad a_3 = 2.746460 \text{ E+00}$$

$$a_1 = 2.350274 \text{ E+00} \quad a_4 = 0.000000 \text{ E+00}$$

$$a_2 = -2.029024 \text{ E+00} \quad t_0 = 0.0000$$



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